Properties of Acids & Bases

Procedure:

- 1. Record data into the table.
- 2. Use the wells in the spot plate to hold the acids and bases.
- 3. Add 5 drops hydrochloric acid to a well in the spot plate, 5 drops of acetic acid into another well, 5 drops of sodium hydroxide into another, and 5 drops ammonium hydroxide into another.
- 4. Dip a glass rod into the hydrochloric acid and touch it to a piece of red litmus. Do the same to the blue litmus and pH paper. Be sure to clean the stir rod. Record your results.
- 5. Into two other wells, add 5 drops of hydrochloric acid. Add 2 drops phenolphthalein into one of the wells and 2 drops universal indicator into another. Record your results.

and 2 drops universal indicator into another. Record your results. 6. Repeat steps 4 & 5 with the acetic acid, sodium hydroxide, and ammonium hydroxide.
7. Add a piece of magnesium to each of the four original wells. Record your results.
8. Place two drops of acetic acid onto your thumb/finger and describe how it feels. Do the same with the
sodium hydroxide. Be sure to wash your hands when finished.
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Questions:
1. List three properties of an acid. List three properties of a base.
2. Describe the feeling of the two solutions.
3. What effect does an acid have on a metal? A base?
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4. What gas was produced in step 7 of the procedure? What test could you do to confirm this gas?
5. In what common household substance might you find hydrochloric acid? Acetic acid? Sodium hydroxide? Ammonium hydroxide?
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- 6. What do the formulas of the acids have in common? The bases?
- 7. Are acids and bases equally dangerous? Explain:

	HCI	HC₂H₃O₂	NaOH	NH₄OH
Blue Litmus				
Red Litmus				
pH paper				
phenolphthalein				
Universal indicator				
magnesium				
Touch test	Do not do this test with HCl			Do not do this test with NH ₄ OH