

## pH and pOH Problems

1) What is the pH of the solution with a hydronium concentration  $[\text{H}_3\text{O}^+]$   $1.47 \times 10^{-4}$ ?

What is the pOH of this solution?

2) What is the pOH of the solution with a hydroxyl concentration  $[\text{OH}^-]$   $2.98 \times 10^{-2}$ ?

What is the hydronium concentration  $[\text{H}_3\text{O}^+]$  of this solution?

3) What is the hydronium concentration  $[\text{H}_3\text{O}^+]$  of a solution with a pH of 7.84?

What is the hydroxyl concentration  $[\text{OH}^-]$  of this solution?

4) What is the hydroxyl concentration  $[\text{OH}^-]$  of a solution with a pH of 3.76?

5) What is the hydronium concentration  $[\text{H}_3\text{O}^+]$  of a solution with a pOH of 2.47?

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