Chapter 6 Preview Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Directions: Write “correct” in the blank following the statement if the statement is true. If the statement is false, cross out the italicized word(s) and write the word(s) in the blank to make the statement true. Also, preceding the statement, write in the page number(s) where this information is found in the text.

**Page**

\_\_\_\_\_1. The sharing of electrons between atoms is called *ionic* bonding. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_2. Electronegativity differences *do not* help in determining the polarity of a bond. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_3. In *a polar* covalent bond, bonding electrons in the molecule are shared equally. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_4. Water has a *molecular formula* of H20, as one oxygen atom is bonded to two hydrogen atoms. \_\_\_\_\_\_

\_\_\_\_\_5.Atoms tend to share or transfer electrons with other atoms until they reach the electron configuration of noble gases. This tendency to have eight electrons in the outer level is called the *octet rule*. \_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_6. Electron-dot diagrams, used to make *Lewis Structures*, show valence electrons in molecules. \_\_\_\_\_\_\_

\_\_\_\_\_7. A pair of electrons that are not part of a bond are called *wasted* pairs. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_8. If two atoms evenly share four electrons, or two pairs, it is a called a *quadruple* bond. \_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_9. There are multiple possible Lewis Structures in a *resonance* structure. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_10. Ionic compounds are made up of positively and negatively charged *isotopes*. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_11. Positive ions tend to join with negative ions forming *electric* bonds. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_12. The three-dimensional, repeating units making up all salts is called the *crystal lattice*. \_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_13. Lattice energy is released in the formation of a *crystal* structure. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_14. Ionic bonds tend to be *weak* bonds. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_15. A *megatomic* ion is a group of two or more bonded atoms that as a whole will bond with other ions to form ionic compounds. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_16. Metallic atoms can be held together by a *sea of electrons* forming metallic bonds. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_17. The valence shell electron pair repulsion theory is used to predict the *size* of a molecule. \_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_18. In order to equalize energy, electrons are sometimes rearranged in a process called *hybridization*. \_\_

\_\_\_\_\_19. Intermolecular forces can cause *strong* molecular bonds. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_20. A molecule containing both positively and negatively charged regions is a *dipole*. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_