Chapter 22 Preview Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Directions: Write “correct” in the blank following the statement if the statement is true. If the statement is false, cross out the italicized word(s) and write the word(s) in the blank to make the statement true. Also, preceding the statement, write in the page number(s) where this information is found in the text.

**Page**

\_\_\_\_\_1. The simplest class of organic compounds are the *carbohydrons*. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_2. A *structural* formula shows the arrangement of atoms in an organic molecule. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_3. Molecules with the same molecular formula but different structure are called *alcohols*. \_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_4. Alkanes are *unsaturated* hydrocarbons. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_5. A hydrocarbon with all single bonds will end with an “*ane*” suffix. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_6. A hydrocarbon with four carbons in its chain will have a “*prop*” prefix. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_7. Using a “yl” suffix in the name of a compound indicates *branching*. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_8. When naming a molecule, we assign numbers to the carbons in the chain to give the *lowest* possible numbers. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_9. Alkanes with 5 to 10 carbons in the chain tend to be *gases*. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_10. Fractional distillation uses *boiling point* differences to separate petroleum into its parts. \_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_11. A double bond in a carbon chain will cause the molecule to be *saturated*. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_12. Alkenes for at least one *triple* bond between carbon atoms. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_13. *Aromatic* compounds have a ring of carbon atoms in the molecule. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_14. A group of atoms that is generally responsible for the characteristics of that organic compound is called a *functional group*. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_15. You can tell by the suffix that propanol is an *aldehyde*. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_16. Hydrogen can easily be replaced by a halogen in a *substitution* reaction. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_17. An amine molecule will have a *neon* atom in its structure. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_18. Water is often the byproduct in *an addition* reaction. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

\_\_\_\_\_19. Monomers build *polymers* in addition reactions. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

\_\_\_\_\_20. The code HDPE on a plastic container tells you it is *polyvinyl chloride*. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.