

Table 3. NAMES AND FORMULAS OF COMMON IONS

Positive	Negative	$C_2H_3O_2^-$
	Acetate	$CH_3COO^-$
Ammonium	$NH_4^+$	Bromide
		$Br^-$
	Carbonate	$CO_3^{--}$
	Hydrogen carbonate (or bicarbonate)	$HCO_3^-$
	Chlorate	$ClO_3^-$
	Chloride	$Cl^-$
	Chlorite	$ClO_2^-$
	Chromate	$CrO_4^{--}$
	Dichromate	$Cr_2O_7^{--}$
	Fluoride	$F^-$
	Hydroxide	$OH^-$
	Hypochlorite	$ClO^-$
	Iodide	$I^-$
	Iodate	$IO_3^-$
	Nitrate	$NO_3^-$
	Nitrite	$NO_2^-$
	Oxalate	$C_2O_4^{--}$
	Hydrogen oxalate (or binoxalate)	$HC_2O_4^-$
	Perchlorate	$ClO_4^-$
	Permanganate	$MnO_4^-$
	Phosphate	$PO_4^{--}$
	Monohydrogen phosphate	$HPO_3^{--}$
	Dihydrogen phosphate	$H_2PO_4^-$
	Silicate	$SiO_3^{--}$
	Sulfate	$SO_4^{--}$
	Hydrogen sulfate (or bisulfate)	$HSO_3^-$
	Sulfide	$S^{--}$
	Hydrogen sulfide (or bisulfide)	$HS^-$
	Sulfite	$SO_3^{--}$
	Hydrogen sulfite (or bisulfite)	$HSO_3^-$
	Tartrate	$C_4H_4O_6^{--}$
	Tetraborate	$B_4O_7^{--}$
	Thiocyanate	$SCN^-$

ACTIVITY SERIES

- Li
- K
- Ca
- Na
- Mg
- Al
- Zn
- Cr
- Fe
- Ni
- Sn
- Pb
- H<sub>2</sub>
- Cu
- Hg
- Ag

Halogens

- F<sub>2</sub>
- Cl<sub>2</sub>
- Br<sub>2</sub>
- I<sub>2</sub>