

Formulas and Nomenclature

I. Name the following compounds:

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| 1. HCl | 19. H ₃ PO ₄ |
| 2. KOH | 20. CsOH |
| 3. HgOH | 21. Li ₂ O |
| 4. KCl | 22. Ca(OH) ₂ |
| 5. FeCl ₃ | 23. CaBr ₂ |
| 6. HNO ₃ | 24. Fe ₂ O ₃ |
| 7. NH ₄ OH | 25. H ₂ SO ₄ |
| 8. Cu ₂ O | 26. FeCO ₃ |
| 9. Al ₂ (SO ₄) ₃ | 27. SO ₃ |
| 10. N ₂ O ₅ | 28. Ba(BrO ₃) ₂ |
| 11. NaOH | 29. Al(OH) ₃ |
| 12. CO ₂ | 30. HClO ₄ |
| 13. HF | 31. NaC ₂ H ₃ O ₂ |
| 14. Pb(OH) ₂ | 32. Na ₂ SO ₃ |
| 15. NH ₄ NO ₃ | 33. H ₂ CO ₃ |
| 16. NaHCO ₃ | 34. HFCO₂ |
| 17. HgO | 35. NH ₄ IO ₃ |
| 18. Zn(NO ₂) ₂ | 36. LiH |

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| 9. manganese dioxide | 29. hydrogen acetate |
| 10. sulfur dioxide | 30. copper (II) nitrite |
| 11. iron (II) sulfate | 31. nitrogen dioxide |
| 12. hypochlorous acid | 32. phosphorus trichloride |
| 13. potassium permanganate | 33. sodium phosphate |
| 14. silver chloride | 34. potassium carbonate |
| 15. copper (II) hydroxide | 35. phosphoric acid = <i>hydrogen phosphate</i> |
| 16. ammonium sulfide | 36. lead (IV) chloride |
| 17. nickel bromide | 37. tin (II) bromide |
| 18. iron (II) oxide | 38. ammonium hydroxide |
| 19. bromic acid = <i>hydrogen bromate</i> | 39. periodic acid |
| 20. ammonium bisulfate | 40. iron (II) hydroxide |
| 21. mercury (I) sulfate | 41. carbon dioxide |
| 22. iron (III) oxide | 42. dinitrogen pentoxide |
| 23. magnesium phosphate | 43. silver oxide |
| 24. nickel bicarbonate | 44. aluminum nitride |
| 25. zinc hydroxide | 45. manganese (II) hydroxide |
| 26. hydriodic acid = <i>hydrogen iodide</i> | 46. ammonium carbonate |
| 27. diphosphorus pentoxide | 47. aluminum oxide |
| 28. aluminum phosphate | 48. antimony pentasulfide |

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