

Directions: Write "correct" in the blank following the statement if the statement is true. If the statement is false, cross out the italicized word(s) and write the word(s) in the blank to make the statement true. Also, preceding the statement, write in the page number(s) where this information is found in the text.

Page

- \_\_\_\_\_ 1. Reactants are changed into products in a *chemical reaction*. \_\_\_\_\_
- \_\_\_\_\_ 2. In a chemical equation *symbols* can be used to show state of matter. \_\_\_\_\_
- \_\_\_\_\_ 3. To satisfy the law of conservation of mass, a reaction must be *balanced*. \_\_\_\_\_
- \_\_\_\_\_ 4. To balance a chemical equation, numbers called *subscripts* are placed in front of the formulas of the elements or compounds. \_\_\_\_\_
- \_\_\_\_\_ 5. Hydrocarbons may react with oxygen to form carbon dioxide and water in a *double displacement* reaction. \_\_\_\_\_
- \_\_\_\_\_ 6. The electrolysis of water is good example of a *synthesis* reaction. \_\_\_\_\_
- \_\_\_\_\_ 7. In a *displacement* reaction, a single element reacts with a compound and displaces another element from the compound. \_\_\_\_\_
- \_\_\_\_\_ 8. Element A cannot displace Element B unless Element A is *lower than* Element B on the activity series. \_\_\_\_\_
- \_\_\_\_\_ 9. The branch of chemistry that deals with quantities of substances in chemical reactions is known as *stoichiometry*. \_\_\_\_\_
- \_\_\_\_\_ 10. The *excess* reactant is the reactant that determines how much product can form in a reaction. \_\_\_\_\_

Directions: Complete the concept map found in question #72 on page 297.

- a. \_\_\_\_\_
- b. \_\_\_\_\_
- c. \_\_\_\_\_
- d. \_\_\_\_\_
- e. \_\_\_\_\_